

"Express Mail" mailing label number EL 594256187US

Date of deposit: November 2, 2001

PATENT
Our Case No. 10225/21 (A18)

IN THE UNITED STATES PATENT AND TRADEMARK OFFICE
APPLICATION FOR UNITED STATES LETTERS PATENT

INVENTORS:

Bruce J. Sabacky
Timothy M. Spittler

TITLE:

PROCESS FOR MAKING
NANO-SIZED STABILIZED ZIRCONIA

ATTORNEY:

G. Peter Nichols
BRINKS HOFER GILSON & LIONE
P.O. BOX 10395
CHICAGO, ILLINOIS 60610

FOR FILING

PROCESS FOR MAKING NANO-SIZED STABILIZED ZIRCONIA

BACKGROUND OF THE INVENTION

[001] The present invention relates to nano-sized stabilized zirconia and, in particular, to an economical method of making nano-sized stabilized zirconia.

[002] Zirconia (ZrO_2) exhibits toughness, wear resistance, hardness, low thermal conductivity and other properties that make it useful in numerous industrial applications.

[003] Stabilized zirconia exhibits high fracture toughness, absorbs energy of impact that shatters other ceramics, and can tolerate thermal gradients better than most other high temperature materials. Because of its high oxygen conductivity at elevated temperatures, stabilized zirconia is also the material of choice for electrolytes in solid-state fuel cells.

[004] There are three commonly occurring and established crystal forms of zirconia: cubic, tetragonal, and monoclinic. The cubic form is the high temperature form and is stable above 2370°C . The tetragonal form is stable between 1170°C and 2370°C . The monoclinic form is stable below 1170°C . The monoclinic to tetragonal phase change is accompanied by a volume change of about 4%. Cooling from the manufacturing temperature often destroys pure zirconia, or gives it inferior mechanical properties. Therefore, it is desirable to stabilize the zirconia in some fashion.

[005] Several methods have been used to form stabilized zirconia. Generally, a stabilizing agent such as calcium, magnesium, yttrium, cerium, or rare earth oxides is added. Zirconia, structurally stabilized by the addition of calcium, yttrium, magnesium, cerium or rare earth oxides has numerous applications. One important application is the manufacture of oxygen sensors. It is also widely used today both for monolithic components and as a coating over refractory metal alloys to serve as a thermal barrier.

[006] Recent advances have shown the advantages of manufacturing nano-sized stabilized zirconias. Nano-sized stabilized zirconia exhibits several very favorable properties, including significant reduction in sintering temperature and ability to deform superplastically under applied stress. Other enhancements include higher diffusivities and possibly higher ionic conductivities. These advantages play heavily both in the

manufacture of solid oxide fuel cells and in producing spray coatings with superior mechanical attributes.

[007] Existing methods to produce nano-sized stabilized zirconium oxide powders include co-precipitation and sol-gel synthesis. Existing gas-phase methods to produce nano-sized stabilized ZrO_2 include inert gas condensation and chemical vapor condensation. These methods, however, are not economical to produce the bulk quantities needed for applications such as thermal barrier coatings and solid oxide fuel cells.

[008] U.S. Pat. No. 6,162,530 discloses a process to make nano-sized powders of ZrO_2 from aqueous solutions. The disclosed process involves atomizing an aqueous solution of the desired metals in a stream of nitrogen and contacting the resulting particles with a spray of a recirculating aqueous solution at controlled pH. The particles produced by this method consist of nanostructured fibers. Further treatment includes sequential heat treatment, ultrasonication, and spray drying.

[009] Although this process may be simpler and cheaper than the co-precipitation or sol-gel synthesis, the process still requires steps that become expensive and difficult to control when extrapolated to commercial production scale, e.g., the efficient mixing of the atomized particles with the recirculating solution requires a special reactor and filtration of the particles will in practice require cumbersome equipment and produce large streams of waste solutions.

[010] The process of the present invention does not produce fibers but particles of approximately the same size in all three dimensions, and solves the above problems by providing an economical and commercially practical method for making stabilized zirconia.

SUMMARY OF THE INVENTION

[011] The present invention provides a process for the manufacture of nano-sized stabilized zirconia that comprises preparing an aqueous feed solution that contains a zirconium salt and a stabilizing agent, converting the feed solution under controlled conditions to form an intermediate, and calcining the intermediate to form agglomerates of nano-sized particles.

[012] The terms “nano-sized” and “nano-size” as used in the following description and claims are meant to refer to particles having linear dimensions in the range from about 1 nm to about 100 nm.

[013] The evaporation is conducted under conditions to achieve substantially total evaporation and to form an intermediate. In particular, the conversion is conducted at a temperature higher than the boiling point of the feed solution but lower than the temperature where significant crystal growth occurs. The intermediate is generally amorphous. Its X-ray diffraction pattern is a very broad, low intensity peak. In a preferred embodiment, the intensity of the X-ray diffraction peak is no more than 15% of the intensity corresponding to a fully formed crystal. In a particularly preferred embodiment, the intermediate is formed as a thin film and preferably has a shape that is spherical or part of a sphere.

[014] The term “substantially total evaporation” or “substantially complete evaporation” refers to evaporation such that the solid intermediate contains less than 15% free water, preferably less than 10% free water, and more preferably less than 1% free water. The term “free water” is understood and means water that is not chemically bound and can be removed by heating at a temperature below 150° C. After substantially total evaporation or substantially complete evaporation, the intermediate product will have no visible moisture present.

[015] The intermediate product is further calcined at a temperature between 400°C and 1300°C and may be milled to liberate the nano-sized particles. The product may be further spray dried and classified to produce agglomerates of nano-sized particles in the desired size range.

BRIEF DESCRIPTION OF THE DRAWINGS

[016] Fig. 1 provides a flow sheet of one embodiment of the process according to the present invention.

[017] Fig. 2 provides a flow sheet to another embodiment of the process according to the invention, where the final product is nano-structured feed material for a thermal spraying process.

[018] Fig. 3 is a scanning electron micrograph of the intermediate product after step 20.

[019] Fig. 4 is a scanning electron micrograph of the product of step 30, calcined at 1000 °C.

[020] Fig. 5 is a scanning electron micrograph of the product of step 30, calcined at 700 °C.

5 **DESCRIPTION OF THE INVENTION**

[021] It has now been found that nano-sized particles of stabilized ZrO_2 can be obtained in an economical manner in a process that includes hydrolysis at a temperature sufficient to insure substantially complete evaporation but low enough to prevent significant crystal growth. The intermediate is generally amorphous. Its X-ray
10 diffraction pattern is a very broad, low-intensity peak. In a preferred embodiment, the intensity of the X-ray diffraction peak is no more than 15% of the intensity corresponding to a fully formed crystal. The process is related to the processes disclosed in U.S. Patent applications Serial Nos. 09/503,365, 09/503,636, and 60/241,041, the entire contents of which are incorporated herein by reference.

[022] Referring to Fig. 1, the general process according to the invention is shown. A starting or aqueous feed solution 10 is comprised of water, an acid, a zirconium salt, and a stabilizing agent. The feed solution is free of insoluble or precipitated material.

[023] The acid may be selected from the group consisting of sulfuric, hydrochloric, and nitric acid. Preferably, the acid is hydrochloric acid. The amount of acid present in the starting solution is at least equal to the amount needed to avoid spontaneous hydrolysis of the zirconium in solution. It may, however, be desirable to provide an excess of acid. For example, when the acid is HCL and the zirconium salt is zirconium oxychloride, the amount of excess acid may vary over a wide range, but is
25 generally less than an amount to provide an excess of about 150 g/l Cl.

[024] The type zirconium salt is not particularly limited so long as it is water-soluble. Accordingly, the zirconium salt includes zirconium oxysulfate, zirconium oxychloride, zirconium oxynitrate, zirconium nitrate, and other water-soluble zirconium salts. Preferably, zirconium oxychloride is used.

[025] The stabilizing agent can be any water soluble element that can impart structural stability to zirconium dioxide. The stabilizing agent is preferably selected from

the group of metals consisting of calcium, magnesium, yttrium, cerium and rare earth oxides. Generally, the metals are provided as metal chlorides or metal oxides.

[026] The starting solution is converted to an amorphous solid intermediate in a conversion process step 20 that involves substantially total, controlled evaporation of the solution and the formation of a thin amorphous film. This process is conducted above the boiling point of the solution but below the temperature where there is significant crystal growth. The water and acid are vaporized and the acid may be recovered by any known process.

[027] The conversion process step is performed in a manner to control the physical form of the amorphous intermediate. Preferably, the conversion process step includes the substantially total evaporation of the solution and hydrolysis of the product resulting from the substantially complete evaporation. The conversion step also preferably successively evaporates part of the feed solution, hydrolyzes the zirconium in solution, and substantially evaporates the remainder of the solution. In a preferred embodiment, the conversion process is conducted using a spray hydrolysis process.

[028] In the preferred spray hydrolysis process, hydrolysis is accomplished by spraying the solution while it is heated at a temperature in the range from about 120° C to about 350° C, preferably in the range from about 200° to about 250° C. Spray hydrolysis may be effected in a spray dryer.

[029] Through control of the operating parameters, including temperature and chemical composition of the starting solution, the characteristics of the solid amorphous intermediate can be reliably controlled within a fairly narrow range. For example, the amorphous intermediate resulting from spray hydrolysis in a spray dryer will be composed of amorphous, hollow, thin-filmed spheres or parts of spheres. The dimensions of the spheres may vary over a wide range, from less than 1 μm to 100 μm in diameter, and the shell thickness in the range from about 30 nm to about 1000 nm. The shells comprise an amorphous, hydrated zirconium oxide containing residual chloride.

[030] Without being bound by any theory, it is believed that spray hydrolysis under the conditions described above produces spheres composed of a thin film of an amorphous solid that can readily be converted to nano-sized stabilized ZrO_2 . Spray

hydrolysis also has the advantage of direct processing of the solution so that the formation of the desired amorphous product and the evaporation of water and HCl are simultaneously accomplished.

[031] The amorphous intermediate zirconia compound resulting from spray hydrolysis is calcined 30 at a temperature and for a length of time sufficient to convert the amorphous zirconia to zirconium dioxide of the desired crystal structure, namely ZrO_2 stabilized in the cubic or tetragonal form by inclusion of yttrium, calcium, magnesium, or rare earth oxides.

[032] Calcination temperatures can range between about 400° and about 1300 C°. Preferably, the calcination is conducted at temperatures ranging from about 600° to 1000° C. The calcination time varies over a wide range, from about 1 h to about 24 h. Lower temperatures will require longer calcination times.

[033] The product of calcination is a thin film having a structure of individual units that can be broken into particles of the desired median size and distribution.

[034] If the starting solution is a chloride solution, any residual chloride in the amorphous zirconia compound resulting from spray hydrolysis is expelled during calcination as HCl gas that can be recovered.

[035] Depending on the intended application, the product of calcination can be further processed by pressing and sintering, for example to make three-dimensional structures, or by milling and particle classification, to produce powders for thermal spraying. Milled particles can also be further processed, for example by tape casting, to make electrolyte structures for solid oxide fuel cells.

[036] Referring to FIG. 2, an alternative embodiment of the present invention is shown. In this embodiment, the product from calcination is milled 40. Milling can be performed by any suitable process such as by milling in water to form an aqueous slurry. The aqueous slurry can then be spray dried 50 to form weakly agglomerated nano-sized stabilized zirconia. Thereafter, the agglomerated particles can be classified 60 according to size with the appropriate size particles being used as thermal spraying powder. Before classifying, the weakly agglomerated particles may be sintered to produce a partially fused agglomerate structure.

[037] The following examples are meant to illustrate the process of the present invention and not to limit the claims.

EXAMPLE I:

5 [038] Reagent grade ZrOCl_2 and YCl_3 are dissolved in a 2 M HCl solution to provide a starting solution containing 39 g/l Zr and 7 g/l Y. The solution is injected through the nozzle of a spray dryer. The spray dryer has a titanium chamber, the air inlet temperature is about 600° C and the outlet is about 250° C. The injection process substantially completely evaporates the solution, hydrolyzes the material, and forms an
10 amorphous intermediate product, which is recovered in a cyclone. The intermediate product is dry and amorphous. The loss on ignition when this product is heated to 800 °C is about 25%. Fig. 3 is an electron micrograph of this intermediate product. This product is further calcined at 1000° C for 6 h. The resulting crystal has 100% cubic structure. Fig. 4 is an electron micrograph of this product, showing part of a spherical shell with a thickness of about 450 nm and individual particles of about 100 nm in size.

EXAMPLE II:

20 [039] Reagent grade ZrOCl_2 and $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$ are dissolved in a 2 M HCl solution to provide a starting solution containing 44 g/l Zr and 1 g/l Ca. The solution is injected through the nozzle of a spray dryer. The spray dryer has a titanium chamber, the air inlet temperature is about 600° C and the outlet is about 250° C. The injection process substantially completely evaporates the solution, hydrolyzes the material, and forms an amorphous intermediate product, which is recovered in a cyclone. The intermediate product is dry and amorphous. The loss on ignition when this product is
25 heated to 800 °C is about 20%. This powder is further calcined at 700° C for 6 h. The resulting crystal has 100% cubic structure. Fig. 5 is an electron micrograph of this product, showing part of a spherical shell with a thickness of about 600 nm and individual particles of about 30 nm in size.

EXAMPLE III:

30 [040] A solution of Zr oxychloride and yttrium chloride is prepared as in example I. The solution is treated in a spray dryer in the same conditions as in example I. The

5

10